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~~(Thermochemistry)~~
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Formulas - Lecture
Review /u0026
Practice ProblemsAP
Chemistry Unit 6

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Review: Thermochemistry

Thermodynamics! AP
Chemistry: 6.1-6.5

Energy Diagrams,
Thermal Equilibrium,
and Heat Capacity

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6.1 (The Nature of Energy) Introduction to chapter 6:

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~~Thermochemistry~~

Calorimetry Concept,

Examples and

Thermochemistry |

How to Pass

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CH7 Atomic Structure

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and Periodicity

Enthalpy of Reaction

Chapter 16

(Spontaneity,

Entropy and Free

Energy) - Part 2 Hess's

Law and Heats of

Formation Section 11

- Thermo Chemistry

(Part 2) Intro to

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~~Tricks to solve~~

~~Thermochemistry~~

~~problems easily |~~

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Enthalpy of
formation
combustion Chapter
7 (Atomic Structure
and Periodicity) - Part
4 Chapter 5 (Gases) -
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Energy Flow and

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Chemical Change

part 1 Chapter 6

thermochemistry 1

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Review Chem-1a AP

~~Chem Thermochem 1~~

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10/19 ~~Chapter 6:~~

~~Thermochemistry~~

~~(37:23)~~ Chapter 6

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162 CHAPTER 6: THERMOCHEMISTRY

To convert the answer to joules, we write: $101.3 \text{ J} / 0.18 \text{ L atm}$
 $1 \text{ L atm} = - \quad \times$
 $= \quad w - 18 \text{ J}$ 6.17 An expansion implies an increase in volume, therefore w must be $- 325 \text{ J}$ (see the defining equation for pressure-volume

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Test Answers

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energy. radiant

energy. thermal

energy. chemical

energy. the capacity

to do work. comes

from the sun and is

earth's primary

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energy source. the energy associated with random motion of atoms and molecules. energy stored within the bonds of chemical substances.

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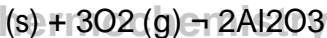
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MULTIPLE CHOICE.

Choose the one
alternative that best
completes the
statement or answers
the question. DH 4AI

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(s) A) exothermic,
released

B) exothermic,
absorbed. AP

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Name_____

MULTIPLE CHOICE.

Choose the one alternative that best completes the statement or answers the question. 1) A

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hermochemistry

that absorbs heat
from the

surroundings is said
to be _____ and has a
_____ ΔH at constant
pressure.

A) endothermic,
positive

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THERMOCHEMISTRY

Substituting into the above equation: $E = 483.6 \times 10^3 \text{ J} - (8.314 \text{ J/mol K})(398 \text{ K})(+1 \text{ mol})$ $E = 4.80 \times 10^5 \text{ J} = 4.80 \times 10^2 \text{ kJ}$

6.28 We initially have 6 moles of gas (3 moles of chlorine and 3 moles of hydrogen). Since our product is 6

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Name _____

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MULTIPLE CHOICE

Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure.

A) endothermic,

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positive. Found: 8
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(Sections 5.1 - 5.4)

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the concepts from

your packet on

Chapter 5

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reserved. 115 14. Two

metals of equal mass

with different heat

capacities are

subjected to the

same amount of heat.

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Which undergoes the smallest change in temperature? a) The metal with the higher heat capacity. b) The metal with the lower heat capacity.

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Oregon State

University 162

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To convert the

answer to joules, we

write: $101.3 \text{ J} / 0.18 \text{ L}$

$\text{atm} \cdot 1 \text{ L} \cdot \text{atm}^{-1} = - \quad \times$

$= \quad w - 18 \text{ J} / 6.17$ An

expansion implies an

increase in volume,

therefore w must be

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-325 J (see the
defining equation for
pressure-volume
work.)

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Ch.6.1: The Nature of
Energy Energy: An
object ' s capacity to
perform work or

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produce heat chemistry

Potential Energy:

Energy due to
position or
composition

(chemical bonds).

Kinetic Energy:

Energy due to the
motion of the object

$\frac{1}{2}mv^2$ KE Law of
Conservation of

Energy: Energy can
neither be created
nor destroyed,

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(Part II) & Chp. 19 -

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(Chp. 1-5) Materials.

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in these relates

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strongly to this chemistry

chapter. Anything that is not included in what we studied you would not be responsible for.

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Mrs. Duffey - FHN

Chapter 6

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w is determined into

or out of the

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Because $E_{univ} = 0$

by the first law of
thermodynamics,

$$E_{sys} = -E_{surr};$$

what happens to the
surroundings is the
exact opposite of
what happens to the
system. chapter 6 test
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