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Reaction and Catalyst RBSE

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explanation □□□□□ □□□ XII

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basics of chemistry. ~the
cell is a complex "chemical
factory. ~it is made up of
atoms, elements, compounds,
and molecules. the atom.
~living and nonliving things
are made up of tiny units.

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Protons have a positive charge; located in the atomic nucleus.

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Composition 1. 100 washers
 $0.110 \text{ g 1 washer} = 11.0 \text{ g}$
(assuming 100 washers is
exact) $100. \text{ g 1 washer } 0.110$
 $\text{g} = 909 \text{ washers}$ 2. $500. \text{ g 1}$
 $\text{cork } 1.63 \text{ g} = 306.7 = 307$
 corks $500. \text{ g 1 stopper } 4.31$
 $\text{g} = 116 \text{ stoppers}$ 1 kg (1000

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g) of corks contains (1000 g
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Analyse the table given
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Title: Chapter 6 Notes The
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6 Notes The Chemistry of Life
. Section 6.1 Atoms and
Their Interactions; 2 Take 5
 is the condition that
results from changes in the

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independent variable. Study
results in an increase in
the amount of living
material and the formation
of new structures. 3 Take 5

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6. (c) 7. (d) 8. (b) 9. (a)

10. (c) NCERT Solutions for
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HOTS Questions. Question 1.
Why standard entropy of an
elementary substance is not
zero whereas the standard

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enthalpy of formation is taken as zero? Answer: A substance has a perfectly ordered arrangement only at absolute zero. Hence, entropy is zero only at ...

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Thermodynamics is the part
of Chemistry which deals

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with exchange of energy in the form of heat and temperature. William Thomson coined this term in 1794 by combining two Greek words Thermes which is heat and Dynamikos meaning powerful.

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Chapter 6 Equilibrium

Chemistry 211 the reaction

we monitor the mass of Ca^{2+}

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remaining in solution and the mass of CaCO_3 that precipitates, the result looks something like Figure 6.1. At the start of the reaction the mass of Ca^{2+} decreases and the mass of CaCO_3 increases. Eventually

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the reaction reaches a point
after which there

Chapter 6

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