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Chapter 12 Solutions review. liquid that dissolves a gas or solid or the solution component in greater amount. when a system in equilibrium is altered by a change of temperature, pressure, or concentration

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variable, the system shifts in equilibrium composition in a way that tends to counteract this change of variable.

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Chapter 12 Review: Solutions Name: Carly 1)

Does a solution have to involve a liquid?

Explain your answer. No; solutions may exist as gases, liquids, or solids. 2) Explain why a suspension is considered a heterogeneous mixture. The particles are so large that they settle out unless the mixture is constantly stirred or agitated.

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Modern Chemistry 5 Solutions CHAPTER 12

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REVIEW Solutions SECTION 2 SHORT ANSWER

Answer the following questions in the space provided. 1. The following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas in the liquid.

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SECTION 1 SHORT ANSWER Answer the following
questions in the space provided. 1. Match the
type of mixture on the left to its
representative particle diameter on the
right. c solutions (a) larger than 1000 nm a

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suspensions (b) 1 nm to 1000 nm b colloids
(c)

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CHAPTER 12 REVIEW Solutions SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Match the type of mixture

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on the left to its representative particle diameter on the right. c solutions (a) larger than 1000 nm a suspensions (b) 1 nm to 1000 nm b colloids (c) smaller than 1 nm 2.

Identify the solvent in each of the following examples:

mc06se cFMsr i-vi

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Chapter 12, Problem Review_and_Discussion 13
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: 5. *How does ...*

Modern Chemistry Chapter 12 Solutions Review
Answers Modern Chemistry Chapter 12-Solutions
Section 1- Types of Mixtures Solutions are
homogeneous mixtures of two or more
substances in a single phase. Soluble
describes a substance as capable of being
dissolved. Solvent is the dissolving medium
in a solution. Chapter 12 Review Solutions
Section 3 Answers

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