

# Read Book Acids Bases And Salts Answers Key

## Acids Bases And Salts Answers Key

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## ~~Bases And Salts Answers~~

Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction. Examples:  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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Question 1. Two solutions A and B have pH values of 5 and 8 respectively. Which solution will be basic in nature ...

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~~Questions with ...~~

Ans:- Salts are classified into four categories on the basis of the strength of acid and base. they are: Salt of strong acid and strong base.e.g NaCl, KCl; Salt of strong acid and weak base. e.g.  $\text{NH}_4\text{Cl}$ ,  $\text{NH}_4\text{I}$ ; Salt of weak acid and strong base. e.g.  $\text{CH}_3\text{COONa}$ ,  $\text{COO}^- \text{Na}^+$  + Salt of weak acid and weak base. e.g.  $\text{CH}_3\text{COO}^- \text{NH}_4^+$  + How is common salt found in nature?

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Bases & Salts. 8.1 Acids, Bases &  
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2. Bases react with . acids to produce  
salts and water; water to produce  
acids and salts; salts to produce acids  
and water; neither acids, salts, nor  
water

~~Acids Bases and Salts Multiple Choice  
Questions and Answers~~

The salts of strong acids and weak  
bases give acidic solution having pH  
less than 7. Example,  $\text{NH}_4\text{Cl}$ ,  
Ammonium Chloride will have pH less  
than 7. The salts of weak acids and  
strong bases give basic solution  
having pH more than 7. Example,  $\text{Na}_2\text{CO}_3$ ,  
Sodium Carbonate will have pH

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more than 7. Question 7.

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Answer A reaction in which an acid and base react with each other to give a salt and water is termed as



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neutralization reaction. For Example: (i)  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$  (ii)  $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$  15. Give two important uses of washing soda and baking soda. Answer Two important uses of washing soda are:

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Science: Ch 2 Acids, Bases ...~~

Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b) Rock salt is the common name for the mineral  $\text{NaCl}$  (halite). Its chemical formula is NaCl. It may be white or light blue or yellow depending upon impurities present in it.

~~Acids Bases and Salts Chapter Wise  
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Acids Bases And Salts Note Taking

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## ~~Acids Bases And Salts Note Taking~~ Answers

(A) inorganic acid (B) organic acid (C)  
soft organic acid (D) strong inorganic  
acid. Answer. Answer: (a) 5. When  
milk of magnesia reacts with acetic  
acid it produces (A) basic salts (B)  
acidic salts (C) neutral salt (D)  
complex salt. Answer. Answer: (c) 6.  
Which of the following phenomena will  
occur when a small amount of acid is

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added to water ...

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Turmeric is a natural indicator. On adding its paste to acid and base separately, which colours would be observed. (a) Yellow in both acid and base (b) Yellow in acid and red in base (c) Pink in acid and yellow in base (d) Red in acid and blue in base. Answer. Answer: (b) Yellow in acid and red in base

## ~~MCQ Questions for Class 7 Science Chapter 5 Acids, Bases ...~~

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with Answers were prepared based on the latest exam pattern. We have Provided Acids, Bases and Salts Class 7 Science MCQs Questions with Answers to help students understand the concept very well. Acids, Bases and Salts Class 7 MCQs Questions with Answers. Choose the correct answer: Question 1.

## ~~MCQ Questions for Class 7 Science Chapter 5 Acids, Bases ...~~

Acids, Bases and Salts In this first quiz on GCSE Chemistry we take a look at one vital part of the subject - acids, alkalis, bases and salts. In it we find out the difference between alkalis and bases, how acids and alkalis react with one another, how neutralisation can produce salts and the relevance of pH values.

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~~GCSE Acids, Bases and Salts | Revise  
pH Values and Alkalis~~

Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties. Acids taste sour and react with metals and polyatomic ions called carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

~~Acids, bases and salts. 7th Grade  
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the concept very well.

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Acid + Base  $\rightarrow$  Salt + Water.

Question 3: Rain becomes acidic due to. carbon dioxide. sulphur dioxide. nitrogen dioxide. all of the above.

Correct Option is : 4. Solution : Rain becomes acidic because of carbon dioxide, sulphur dioxide and nitrogen dioxide. They dissolve in rain drops to form carbonic acid, sulphuric acid and nitric acid ...

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